WYDZIAŁ INŻYNIERII
LĄDOWEJ I TRANSPORTU

POLITECHNIKA POZNAŃSKA

POZNAN UNIVERSITY OF TECHNOLOGY
INSTITUTE OF BUILDING ENGINEERING
DIVISION OF BUILDING AND BUILDING MATERIALS

BUILDING CHEMISTRY

LAB 2

FOUNDATIONS OF QUANTITATIVE CHEMICAL ANALYSIS

DETERMINATION OF SODIUM HYDROXIDE CONCENTRATION

WITH ACID–BASE TITRATION METHOD

Names…………………………………………………………………………….

Group…………………………………………… | Date…………………………….....

1. Procedure

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2. Test results

Note the chemical equations of performed reaction.

CALCULATION

Molar concentration of the NaOH solution:

$C = V_{mv} \cdot 0,1 \ (mol/dm^3)/V_{NaOH}$

Percent concentration of the NaOH solution:

$$C_p = \frac{V_{mv} \cdot 0,1 \ (mol/dm^3) \cdot 40 \ (g/mol)}{1000 \ (g/dm^3) \cdot V_{NaOH}} * 100\%$$

Weight of NaOH:

$m_{NaOH} = V_{mv} \cdot 0,1 \ (mol/dm^3) \cdot 40 \ (g/mol)$

Nº	Titrant volume (HCl 1 mol/dm³) used in analysis [cm³]	Analyte volume (NaOH)
1	$V_1=$	$V_{NaOH}=$
2	$V_2=$	$C=$
3	$V_3=$	$m_{NaOH}=$
	$V_{mv}=$	$C_p=$

3. Conclusions

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